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diprotic acid. A student determined the value of x in the
formula of hydrated ethanedioic acid,
 $\text{HOOC} - \text{COOH} \cdot x\text{H}_2\text{O}$, by titrating a known mass of the
acid with a $0.100 \text{ mol dm}^{-3}$ solution of $\text{NaOH} (\text{aq})$.
 0.795 g of ethanedioic acid was dissolved in distilled

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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.

$C_2H_4(g)$	+	$H_2(g)$	\rightarrow	$C_2H_6(g)$	3
0 2		2		$CO_2(g)$	+ 3
				H_2O	- 1 4 1 1
					k J m o l - 1
					- 1 - 2 8 6
					k J
					m o l - 1
					- 1 - 1 5 6 0
					k J m o l - 1

(hydrogenation) (l)
Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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